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Experiment 1-

Topic

Volmetric analysis - Stoichiometry. D1

Aim

To determine the exact concentration of a monobasic acid, HX. D2

Apparatus

~~Burette~~ Burette, pipette, retort stand, conical flask, dropper, beaker, volumetric flask.

Materials

Sodium hydroxide (NaOH), phenolphthalein indicator, monobasic acid (HX), distilled water. D3

Introduction

Neutralisation occur when an acid react with base. In laboratory, the concentration of an unknown acid can be determined using titration technique. Therefore, this experiment is conducted to determine the unknown acid concentration.

Procedure

a) Preparing sodium hydroxide, NaOH: D4

$$i. \frac{mV}{1000} = \frac{(250)(0.5)}{1000}$$

$$= 0.125$$

$$\therefore \text{mass of NaOH} : 0.125 \times (23 + 16 + 1)$$

$$= 5g$$

1) 5g of NaOH pellet is dissolve using distilled water in a beaker to prepare 250 ml of 0.5 mol dm^{-3} of NaOH.

2) The solution is transferred into 250 ml volumetric flask. The beaker is rinsed with distilled water and pour into the volumetric flask.

3) More distilled water is added to volumetric flask until it reached the marks. The flask is shaken well.

No.

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